



**BISHOP COTTON WOMEN'S CHRISTIAN COLLEGE**  
C.S.I Karnataka Central Diocese  
#19,3<sup>rd</sup> Cross, C.S.I Compound, Mission Road, Bengaluru – 560027  
Affiliated to Bengaluru City University

Contact No: 080 – 22212933/22129880  
Email: [principal@bcwcc.edu.in](mailto:principal@bcwcc.edu.in) Website: [www.bcwcc.edu.in](http://www.bcwcc.edu.in)

## **DEPARTMENT OF CHEMISTRY**

### **PROGRAM OUTCOMES**

By the end of the program the students will be able to:

(Refer to literature on outcome based education (OBE) for details on Program Outcomes)

1. **PO. 1:** To create enthusiasm among students for Analytical chemistry and its application in various fields of life.
2. **PO. 2:** To provide students with broad and balanced knowledge and understanding of key concepts in Analytical chemistry
3. **PO. 3:** To develop in students a range of practical skills so that they can understand and assess risks and work safely measures to be followed in the laboratory.
4. **PO. 4:** To develop in students the ability to apply standard methodology to the solution of problems in chemistry
5. **PO. 5:** To provide students with knowledge and skill towards employment or higher education in chemistry or multi-disciplinary areas involving chemistry.
6. **PO. 6:** To provide students with the ability to plan and carry out experiments independently and assess the significance of outcomes and to cater to the demands of chemical Industries of well-trained graduates
7. **PO. 7:** To develop in students the ability to adapt and apply methodology to the solution of unfamiliar types of problems.
8. **PO. 8:** To instil critical awareness of advances at the forefront of chemical sciences, to prepare students effectively for professional employment or research degrees in chemical sciences and to develop an independent and responsible work ethics.



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**DEPARTMENT OF CHEMISTRY**  
**COURSE OUTCOMES**

**I SEM: ANALYTICAL/INORGANIC AND ORGANIC CHEMISTRY**

At the end of the course the student should be able to:

*(Write 3-7 course outcomes. Course outcomes are statements of observable student actions that serve as evidence of knowledge, skills and values acquired in this course)*

1. Explain basic laboratory practices like calibration of glassware, sampling, handling acids and safety precautions.
2. Prepare the solutions after calculating the required quantity of salts in preparing the reagents/solutions and dilution of stock solution.
3. Describe the limitations of Classical Mechanics which necessitated the development of Quantum Mechanics.
4. Solve the Schrodinger's equation to obtain wave function for a basic type of Potential in one dimension and predict the shapes of orbitals as well as probability Distributions
5. To justify the need for quantum mechanical structure of atoms
6. Describe the periodicity in physical and chemical properties. Of elements in the Periodic table.
7. Explain the nature of bonding in organic compounds using concepts such as Conjugation, resonance, etc.
8. Learn methods of syntheses of alkanes, alkenes and alkynes along with their

**II SEM: Analytical and Organic Chemistry-II**

- 1) Interrelationship among frequency, wavelength and wave number and importance of validation parameters of an instrumental method will be taught
- 2) Principle, instrumentation and applications of spectrophotometry, nephelometry and turbidometry will be taught
- 3) Fundamentals of separation methods and principles of paper, thin layer and column chromatography will be taught
- 4) Principle, types and applications of solvent extraction will be taught
- 5) Principle and mechanism of ion-exchange, types of resins and domestic and industrial applications of ion-exchange chromatography will be taught
- 6) The concept of mechanism and its importance will be taught to the student



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- 7) Concept and importance of intermediates in organic chemistry will be taught taking proper examples
- 8) The various techniques for identification of reaction mechanism will be taught to the student taking proper examples
- 9) Concept of stereochemistry and its importance will be taught.
- 10) The various projection formulae and the techniques of designating the molecules into R, S, D, L will be taught taking proper examples.
- 11) The theory and concept of Cis-, Trans- isomerism and its importance and the techniques to differentiate between them will be taught taking examples

### III SEM: PRACTICALS

#### Course Objectives

- 1) To impart skills related to preparation of stock and working solutions and handling of instrumental methods
- 2) To know the principle of colorimetric analysis and construction of calibration plot
- 3) To understand the chemistry involved in colorimetric determination of metal ions and anions
- 4) To determine R<sub>f</sub> values of different metal ions present in a mixture
- 5) To impart knowledge on the importance of functional groups in organic compounds.
- 6) Techniques to identify the functional groups in an compound by performing physical and chemical tests
- 7) To record its melting point/boiling point.
- 8) To prepare suitable derivative for that compound and to characterize it.

### SEM IV: Inorganic and Physical Chemistry-II

#### Course Objectives:

Students learn about

1. Different types of bonding in molecules/compounds/ions
2. The structures of molecules/compounds/ions based on different models/theories
3. Properties of compounds based on bonding and structure



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4. The fundamentals of thermodynamics including the laws, the concept of entropy and free energy functions and their applications.
5. The concepts of surface chemistry, catalysis and their applications.
6. The theoretical and experimental aspects of chemical kinetics including basic theories of reaction rates and methods of determining order.
7. Electrochemistry dealing with electrolytes in solution. Conductance measurements and applications. Concept of ionic mobility and their determination.

### **Course outcomes:**

After the completion of this course, the student would be able to

1. Predict the nature of the bond formed between different elements
2. Identify the possible type of arrangements of ions in ionic compounds
3. Write Born - Haber cycle for different ionic compounds
4. Relate different energy parameters like, lattice energy, entropy, enthalpy and solvation energy in the dissolution of ionic solids
5. Explain covalent nature in ionic compounds
6. Write the M.O. energy diagrams for simple molecules
7. Differentiate bonding in metals from their compounds
8. Learn important laws of thermodynamics and their applications to various thermodynamic systems
9. Understand adsorption processes and their mechanisms and the function and purpose of a catalyst
10. Apply adsorption as a versatile method for waste water purification.
11. Understand the concept of rate of a chemical reaction, integrated rate equations, energy of activation and determination of order of a reaction based on experimental data
12. Know different types of electrolytes, usefulness of conductance and ionic mobility measurements.
13. Determine the transport numbers

### **IV SEM: PRACTICALS**

At the end of the course student would be able to



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1. Understand the chemical reactions involved in the detection of cations and anions.
2. Explain basic principles involved in classification of ions into groups in semi-micro qualitative analysis of salt mixture
3. Carryout the separation of cations into groups and understand the concept of common ion effect.
4. Understand the choice of group reagents used in the analysis.
5. Analyse a simple inorganic salt mixture containing two anions and cations
6. Use instruments like conductivity meter to obtain various physicochemical parameters.
7. Apply the theory about chemical kinetics and determine the velocity constants of various reactions.
8. Learn about the reaction mechanisms.
9. Interpret the behaviour of interfaces, the phenomena of physisorption and chemisorptions and their applications in chemical and industrial processes.
10. Learn to fit experimental data with theoretical models and interpret.